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finity for each other is weak. When a stream of nitrous gas is passed into this compound gas, it unites with the oxygen and forms nitrous acid, leaving the chlorine gas unaltered in its properties.

The effect of oxymuriatic acid in bleaching he explained by the affinity of this gas for the hydrogen of water, forming with it muriatic acid, which acted on the vegetable fibre; the oxygen of the water acting at the same time on the colouring matter. The corrosive effects of muriatic acid during this process are lessened if the oxymuriat of lime be used. The oxymuriat of potash is the least prejudicial, but its price will prevent its application to the purpose of bleaching. Dr. Davy stated, he had found that the oxymuriat of magnesia may be used with great advantage; though its bleaching property is not so rapid in its operatioin, it is much less injurious than oxymuriat of lime which is commonly used. The oxymuriat of magnesia will serve repeatedly for the same purpose. If heat be applied to it after it has been used, the hydrogen is expelled, and it is restored to its former state

[*London Monthly Magazine.*]

Remarks on the subject of bleaching would be very acceptable from our chemical readers. In this country, notwithstanding the expense is greater, muriat of potass is much used instead of this muriat of lime, on account of its greater safety, in case sufficient washing is not afterwards applied. In this point, there is always a considerable risque, as perhaps in none of the processes of bleaching are we more deficient, than in washing. To a neglect in this respect we are inclined to attribute most of the damages arising in bleaching in this country.

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